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## A gold-nanoparticle stoppered [2]rotaxane

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The construction of molecular machines has captured the imagination of scientist for decades. Despite significant progress in the synthesis and studies of the properties of small-molecule components (smaller than 2-5 kilo Daltons) challenges regarding incorporation of molecular components into real devices are still eminent. Nano-sized molecular machines operate the complex biological machinery of life, and the idea of mimicking the amazing functions using artificial nano-structures is intriguing. Both in small-molecule molecular machine components and in many naturally occurring molecular machines, mechanically interlocked molecules and structures are key functional components. In this work we describe our initial efforts to interface mechanically-interlocked molecules and gold-nanoparticles (AuNPs); the molecular wire connecting the AuNPs is covered in an insulating rotaxane-layer, thus mimicking the macroscopic design of a copper wire. Taking advantage of recent progress in the preparation of supramolecular complexes of the cucurbit[7]uril (CB[7]) macrocycle, we have prepared a bis-thiol functionalised pseudo-rotaxane that enables us to prepare a AuNP-stoppered [2]rotaxane in water. The pseudo-rotaxane is held together extremely tightly ( $K_a > 10^{13} \text{ M}^{-1}$ ),  $K_a$  being the association constant. We have studied the solution and gas phase guest-host chemistry using NMR spectroscopy, mass spectroscopy, and electrochemistry. The bis-thiol functionalised pseudo-rotaxane holds, further a ferrocene unit in the centre of the rotaxane; this ferrocene unit enables us to address the system in detail with and without CB[7] and AuNPs using electrochemical methods.

### Introduction

Mechanically-interlocked molecules are of significant importance for the development of molecular machines at the molecular level.<sup>1-3</sup> At the small-molecule level the work of Sauvage, Sanders, Leigh, Stoddart, and others since the 1970's has enabled the preparation of complex structures ranging from rotaxanes through catenanes to molecular knots.<sup>1-4</sup> Exploitation of the unique properties of these structures has enabled complex operations such as unidirectional motion and on-off-switches for electronic components.<sup>1-5</sup> The importance of naturally occurring systems that feature mechanically-interlocked structures are being realised in these years. The discovery of knotted DNA strands and lasso peptides are recent examples.<sup>6, 7</sup>

Nano-engineering offers unique opportunity for bridging the gap between the many-nanometre size-regimes of the biological world and the Ångström world of small molecules. Artificial molecular machine immobilized surfaces (AMMISs),

where for example (pseudo)rotaxanes have been incorporated onto gold surfaces, are examples of this.<sup>8-10</sup> The preparation of metallic nanoparticles (NPs) with well controlled size, multifarious functional groups and manipulation of their properties has revolutionized nanoscience and nanotechnology. The most studied metal NPs are gold nanoparticles (AuNPs), now a core nanoscale physical entity with applications ranging from molecular electronics and molecular biology to self-assembly of dynamic combinatorial libraries that lead to sophisticated molecular recognition phenomena.<sup>11-13</sup> In the context of molecular electronics, AuNPs can be regarded as potential ultra-small electrodes, and the unique optical and electronic properties of AuNPs have been explored for development of biosensors and biomedicine.<sup>14-23</sup>

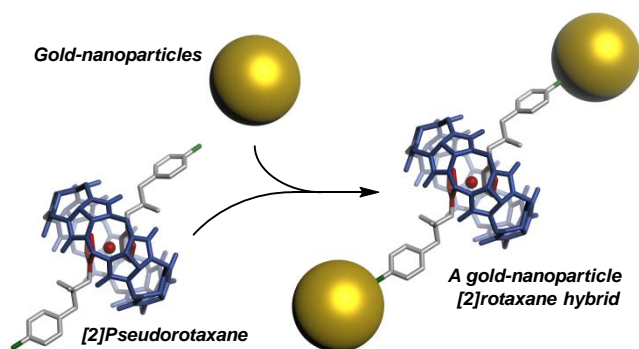
Connecting AuNPs with conducting molecular wires or electrochemically active molecular components is an intriguing thought. If AuNPs are connected via a mechanically interlocked structure, for example a pseudo-rotaxane, one can envisage a molecular wire connecting the AuNPs covered in an insulating layer, thus mimicking the macroscopic design of a copper wire. Success in this area would potentially enable constructing very small electronic components (Fig. 1).

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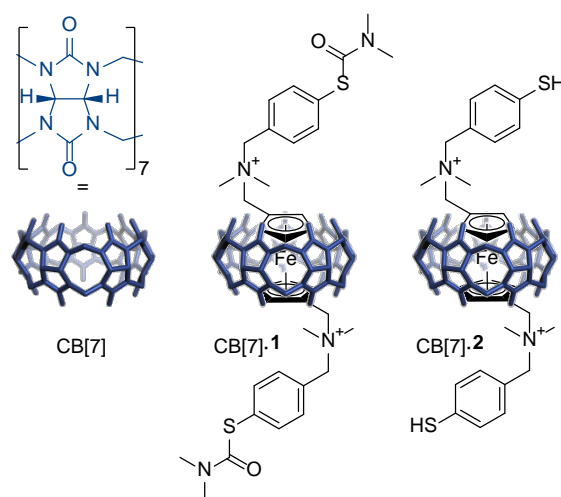
† Electronic Supplementary Information (ESI) available: Experimental procedures, spectral characterisation data, details of the binding studies, and a description of the electrochemical measurements. See DOI: 10.1039/x0xx00000x



**Fig. 1** Schematic representation of the formation of a gold-nanoparticle [2]rotaxane hybrid from a [2]pseudorotaxane and gold-nanoparticles.

Recent developments in the area of supramolecular chemistry of the macrocycles cucurbit[*n*]urils (CB[*n*]) in aqueous solution has been utilized to enable formation of very stable pseudo-rotaxanes; in 2007 a bis-dialkylammonium ferrocene compound with an association constant to CB[7] of  $3 \cdot 10^{15} \text{ M}^{-1}$  in water was discovered, and in 2014 the first host-guest pair with a dissociation constant as small as  $10^{-18} \text{ M}$  was reported for CB[7].<sup>24</sup> Such high binding affinities mean that the complexes are essentially completely intact, even at low concentrations, which makes these pseudo-rotaxanes compatible with the concentrations where AuNPs are prepared.

We report here the preparation of a [2]rotaxane with AuNPs functioning as stoppers and an electroactive ferrocene core structure surrounded by a CB[7] moiety. To achieve this, the preparation of a pseudo-rotaxane with two thiols pointing out of a CB[7] is essential. Specifically, we describe the synthesis and investigation of two inclusion complexes that are based on the CB[7] host and a novel ferrocene based bis-ammonium cationic guest molecules functionalised with thiols (Chart 1). The complexes have been systematically characterised by liquid chromatography–mass spectrometry (LC-MS) analysis, 1D and 2D nuclear magnetic resonance (NMR) spectroscopy as well as electrochemical investigations. The electrochemical data are of particular interest as they disclose detailed insight into the dynamics of the pseudo-rotaxane structure. <sup>1</sup>H NMR competition experiments reveal strong binding affinities of the complexes ( $K_a \sim 10^{13} \text{ M}^{-1}$ ) and the possibility of reversibly removing the functionalised guest upon treatment with a competitive guest molecule. Finally, we confirm the formation of a simple [2]rotaxane structure by transmission electron microscopy (TEM).



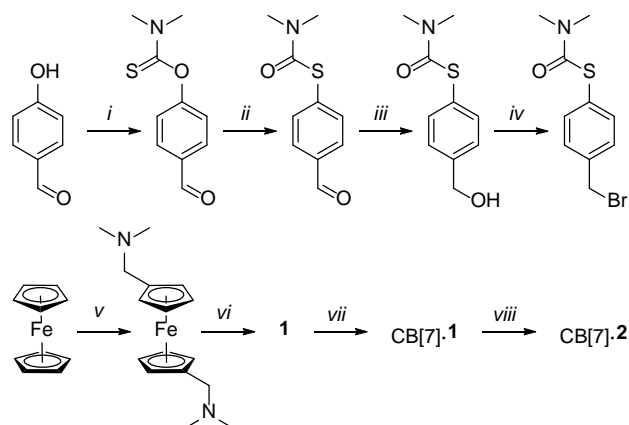
**Chart 1** Structures of the CB[7] host and the inclusion complexes between CB[7] and the ferrocene guests **1** and **2** reported in this work.

## Results and discussion

### Synthesis of the inclusion complexes

The guest **1** was synthesised using a convergent synthesis strategy (Scheme 1). The phenyl moiety was introduced from 4-hydroxybenzaldehyde and the *O*-thiocarbamate formed through a reaction with *N,N*-dimethylthiocarbamoyl chloride. A thermally induced Newman-Kwart rearrangement<sup>25</sup> converted the *O*-thiocarbamate into the corresponding *S*-thiocarbamate. Selective reduction using NaBH<sub>4</sub> transformed the aldehyde into the alcohol and was followed by reaction with PBr<sub>3</sub> to generate the benzyl bromide intermediate.

The dialkylation of the ferrocene moiety was conducted through lithiation of ferrocene with *n*-BuLi followed by *in situ* reaction with Eschenmoser's salt.<sup>26</sup> Adding the benzyl bromide intermediate to the alkylated ferrocene intermediate formed the quaternary ammonium cations via a Menshutkin reaction yielding the bromide salt of guest **1**.



**Scheme 1** Synthesis of **1**, CB[7]·**1**, and CB[7]·**2**. Reagents and conditions: (i) DABCO, (CH<sub>3</sub>)<sub>2</sub>NCSCl, NMP, 50 °C, 17 h, 95 %; (ii) Ph<sub>2</sub>O, 270 °C, 45 min, 79 %; (iii) NaBH<sub>4</sub>, MeOH, 0 – 25 °C, 21 h, 97 %; (iv) PBr<sub>3</sub>, CH<sub>2</sub>Cl<sub>2</sub>, 0 °C – reflux, 21 h, 78 %; (v) (a) *n*-BuLi, TMEDA, hexane, 25 °C, 22 h, (b) [(CH<sub>3</sub>)<sub>2</sub>N(CH<sub>3</sub>)<sub>2</sub>]<sup>+</sup>, 25 °C – reflux, 22 h, 69 %; (vi) benzyl bromide, MeOH, reflux, 10 min, 78 %; (vii) CB[7], KCl, H<sub>2</sub>O, 25 °C, spontaneously, quant.; (viii) NaOH, KCl, H<sub>2</sub>O, 25 °C, 22 h, 82 %.

The inclusion complex CB[7]·1 was formed spontaneously when mixing equimolar amounts of guest 1 together with the CB[7] host, which was synthesised according to a procedure reported previously.<sup>27</sup> Subsequent basic hydrolysis of the *S*-thiocarbamates in CB[7]·1 gave the CB[7]·2 complex. All compounds were characterised by <sup>1</sup>H and <sup>13</sup>C NMR spectroscopy, high-resolution ESI mass spectrometry, and elemental analysis thus confirming their identities (see Supporting Information).

### Characterisation by NMR spectroscopy

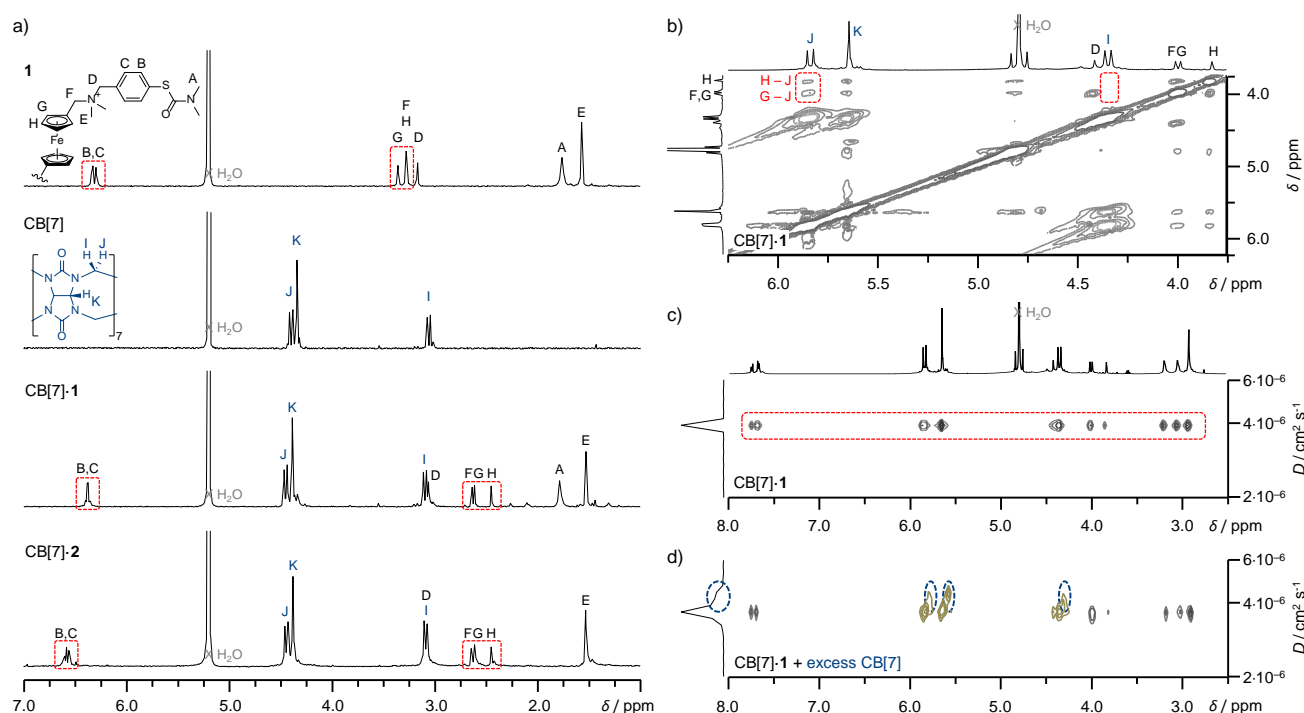
<sup>1</sup>H NMR spectroscopic analysis provides an initial assessment of the binding interactions between the cationic ferrocene guests and the CB[7] host. The <sup>1</sup>H NMR spectra of 1, CB[7], and CB[7]·1 are shown in Fig. 2a. The complex formation induces changes in the chemical shift values of both CB[7] and 1. The presence of CB[7] leads to a pronounced upfield shift of the proton resonances in the cyclopentadienyl (Cp) ring with  $\Delta\delta$  values of 0.75 ppm for *H*<sub>G</sub> and 0.83 ppm for *H*<sub>H</sub>. In contrast, the signals corresponding to the protons on the substituted phenyl ring (*H*<sub>A</sub>, *H*<sub>B</sub>, and *H*<sub>C</sub>) experience only a small downfield shift. The reverse trend in the two moieties of 1 indicates entirely different environments for the ferrocene and phenyl group, showing how the ferrocene residue is positioned inside the cavity of CB[7] upon complexation, while the phenyl moiety remains outside. The methylene and *N*-methyl protons (*H*<sub>D</sub> and *H*<sub>E</sub>) are less shielded, moving upfield with  $\Delta\delta$  values of 0.11 and 0.04 ppm, respectively, which indicates that they are closer to

the cavity than the protons *H*<sub>A</sub> – *H*<sub>C</sub>. Finally, the methylene protons (*H*<sub>F</sub>) that bridge the ferrocene unit to the positively charged nitrogen are shifted upfield ( $\Delta\delta$  0.64 ppm) along with the ferrocene protons, providing clear evidence that they are all fully included by CB[7]. Similar spectral changes were observed for CB[7]·2 (Fig. 2a, bottom).

A 2D ROESY experiment on the complex CB[7]·1 further supports that the ferrocene moiety in 1 is located inside the cavity of CB[7]. The <sup>1</sup>H-<sup>1</sup>H ROESY spectrum displays strong electronic couplings shown as NOE cross peaks from *H*<sub>G</sub> and *H*<sub>H</sub> on ferrocene to CB[7] proton *H*<sub>I</sub> and none to *H*<sub>J</sub> (Fig. 2b). This suggests that the ferrocene moiety and the cavity of CB[7] are in close proximity through space thus forming the inclusion complex CB[7]·1.

Diffusion ordered <sup>1</sup>H NMR spectroscopy (DOSY) on CB[7]·1 was consistent with the presence of a single species in solution with a diffusion coefficient (*D*) of  $3.8 \cdot 10^{-6} \text{ cm}^2 \text{ s}^{-1}$  (Fig. 2c). This result points towards a high association constant between CB[7] and 1, as the DOSY peaks for these two species converge, showing how the host and guest diffuse together as a single molecular entity. Adding excess CB[7] to this sample resulted in two sets of separated signals with different diffusion coefficients; one corresponding to CB[7]·1 with a lower *D* value and one with a higher value corresponding to the smaller CB[7] species ( $D_{\text{CB}[7]} = 4.2 \cdot 10^{-6} \text{ cm}^2 \text{ s}^{-1}$ ; Fig. 2d).

The NMR spectroscopic data thus provide strong evidence for the location of the CB[7] host around the ferrocene residue in the CB[7]·1 and CB[7]·2 complexes.



**Fig. 2** NMR analysis of the compounds synthesised showing a) stacked <sup>1</sup>H NMR spectra (100 mM KCl in 4 M DCl<sub>(aq)</sub>, 500 MHz, 25 °C) of 1, CB[7], CB[7]·1, and CB[7]·2; b) partial <sup>1</sup>H-<sup>1</sup>H ROESY spectrum (500 MHz, 100 mM KCl in D<sub>2</sub>O, 25 °C) of CB[7]·1; c) <sup>1</sup>H DOSY spectrum (500 MHz, 100 mM KCl in D<sub>2</sub>O, 25 °C) of CB[7]·1 ( $D_{\text{CB}[7] \cdot 1} = 3.8 \cdot 10^{-6} \text{ cm}^2 \text{ s}^{-1}$ ); and d) <sup>1</sup>H DOSY spectrum (500 MHz, 100 mM KCl in D<sub>2</sub>O, 25 °C) of CB[7]·1 added excess CB[7] where the blue signals arise from excess CB[7] ( $D_{\text{CB}[7]} = 4.2 \cdot 10^{-6} \text{ cm}^2 \text{ s}^{-1}$ ). Assignment of the signals is based on the labelling shown on the structures in the figure.

### Association constants

In the preliminary  $^1\text{H}$  NMR experiments it was observed that the  $\text{CB}[7]\cdot\mathbf{1}$  and  $\text{CB}[7]\cdot\mathbf{2}$  complexes formed quantitatively with only millimolar concentrations of host and guest, suggesting that the association constant ( $K_a$ ) is high. Further,  $^1\text{H}$  NMR spectra of  $\text{CB}[7]\cdot\mathbf{1}$  when an excess amount of guest was present showed, how the signals of the free and bound guest were simultaneously observed, revealing that the rate of exchange between these species is slow on the NMR time scale. These findings were anticipated since similar observations in other  $\text{CB}[7]$ -ferrocene complexes have been reported, where the high stability was proven through extremely high binding affinities ( $10^9 - 10^{15} \text{ m}^{-1}$ ).<sup>24, 28, 29</sup>

The large association constant precludes determination of  $K_a$  for the complex formation between  $\text{CB}[7]$  and  $\mathbf{1}$  directly by NMR spectroscopy as they exceed the experimentally accessible range ( $10 < K_a < 10^4 \text{ m}^{-1}$ ).<sup>30</sup> Instead, a competition  $^1\text{H}$  NMR spectroscopic experiment was carried out, where guest  $\mathbf{1}$  together with a reference guest (1,1'-bis(*N,N,N*-trimethylammoniomethyl)ferrocene,  $\mathbf{3}$ )<sup>29</sup> was allowed to compete for a limited amount of the host,  $\text{CB}[7]$ . With this procedure an association constant of  $1.1 \cdot 10^{13} \text{ M}^{-1}$  for the complexation between  $\text{CB}[7]$  and  $\mathbf{1}$  in 100 mM KCl in  $\text{D}_2\text{O}$  was determined (Supporting Information, Section S3).

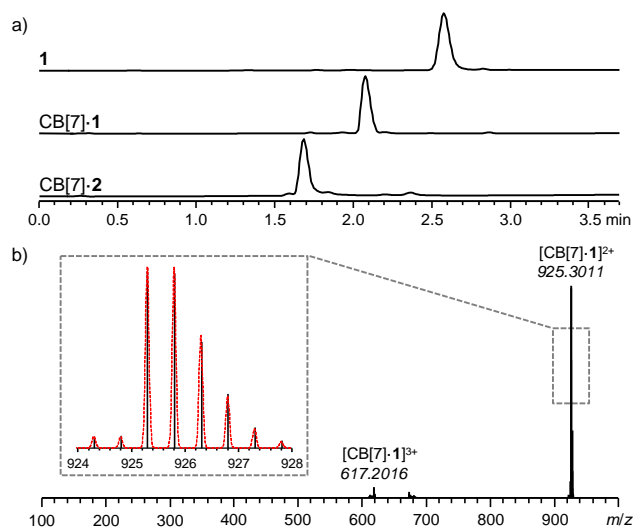
The  $K_a$  value obtained for  $\text{CB}[7]\cdot\mathbf{1}$  is one of the highest binding affinities reported for a monovalent synthetic host-guest complex and is comparable in strength to the in supramolecular chemistry and biology widely used Avidin-D-biotin pair.<sup>31</sup> This result represents the unique binding properties of  $\text{CB}[7]$  towards the guest  $\mathbf{1}$  and shows that high binding affinities competitive with biologically receptor molecules can be achieved for purely synthetic systems.

Our attempts to obtain single crystals of  $\text{CB}[7]\cdot\mathbf{1}$  and  $\text{CB}[7]\cdot\mathbf{2}$  suitable for X-ray diffraction analysis failed, verifying the difficulties often encountered when trying to crystallize  $\text{CB}[7]$ -guest complexes.<sup>32</sup> However, the crystal structure of the complex  $\text{CB}[7]\cdot\mathbf{3}$  previously obtained can be used as a model to describe the core interactions in  $\text{CB}[7]\cdot\mathbf{1}$ .<sup>24</sup> Here, the X-ray structure shows how the ferrocene residue is included by the  $\text{CB}[7]$  cavity filling approximately 55 % of the host cavity volume, which is the optimum cavity-filling fraction as proposed by Mecozzi and Rebek.<sup>33</sup> The near perfect size/shape complementarity between the rigid  $\text{CB}[7]$  cavity and the ferrocene moiety is another factor leading to the formation of the highly stable  $\text{CB}[7]\cdot\mathbf{1}$  complex.

### Characterisation by LC/MS

Mass spectrometry is now an important tool for characterisation of stable inclusion complexes in the gas phase. The technique can provide unique insight into the formation of the complexes, their stoichiometries, and thereby serve as a supplement to condensed phase studies.<sup>34</sup> A combination of high performance liquid chromatography (HPLC) and high-resolution electrospray ionization time-of-flight mass spectrometry (HR-ESI-TOF-MS) was employed for the studies of the inclusion complexes  $\text{CB}[7]\cdot\mathbf{1}$  and  $\text{CB}[7]\cdot\mathbf{2}$ . Fig. 3a shows the

HPLC chromatograms of  $\mathbf{1}$ ,  $\text{CB}[7]\cdot\mathbf{1}$ , and  $\text{CB}[7]\cdot\mathbf{2}$  monitored at 255 nm where  $\text{CB}[7]$  alone is not UV active. Each chromatogram presents one sharp, well-defined peak with a retention time of 2.6, 2.1, and 1.7 minutes as the analyte changes from  $\mathbf{1}$  through  $\text{CB}[7]\cdot\mathbf{1}$  to  $\text{CB}[7]\cdot\mathbf{2}$ . The distinct peaks indicate that the inclusion complexes persist throughout the column and do not degrade or dissociate.



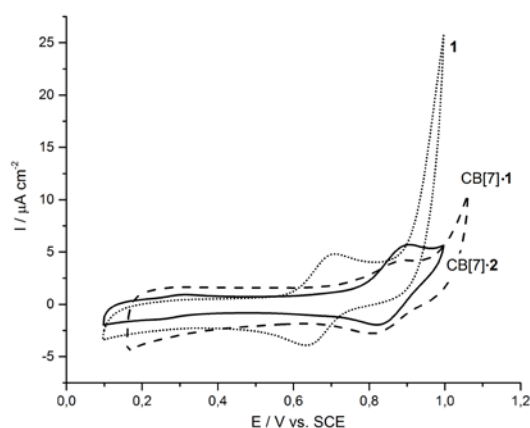
**Fig. 3** LC/MS analysis of the target compounds showing a) stacked HPLC chromatograms (255 nm) of  $\mathbf{1}$  (top),  $\text{CB}[7]\cdot\mathbf{1}$  (middle), and  $\text{CB}[7]\cdot\mathbf{2}$  (bottom), and b) HR-ESI-TOF-MS spectrum of  $\text{CB}[7]\cdot\mathbf{1}$ . The inset shows a zoom of the isotopic distribution of the main peak where the calculated distribution is shown as red dotted lines while the black lines are the measured spectrum.

The corresponding mass spectrum of  $\text{CB}[7]\cdot\mathbf{1}$  is shown in Fig. 3b where the main peak in the spectrum, at  $m/z$  ratio of 925.3011, is a doubly charged peak shown in the inset together with the calculated isotopic distribution. This peak is consistent with one doubly charged bis-ammonium cation  $\mathbf{1}$  enclosed by the neutral  $\text{CB}[7]$  host,  $[\text{CB}[7]\cdot\mathbf{1}]^{2+}$ , verifying the 1:1 inclusion complex. A second minor peak, at  $m/z$  ratio of 617.2016, which match the triply charged inclusion complex  $[\text{CB}[7]\cdot\mathbf{1}]^{3+}$  is also present. We suggest that the third positive charge is acquired through oxidation of the ferrocene unit in the compound thereby reflecting the relative ease of oxidation of the ferrocene/ferrocenium unit in the compound (see the electrochemical characterisation section below), as seen previously for related compounds.<sup>35</sup> Non-covalent complexes tend to dissociate during LC/MS runs. This can happen both when the complex passes through the polar, aqueous mobile phase in the column of the HPLC or in the spray process in the mass spectrometer, where the association energy of the complex becomes lower than the dissociation energy. However, here the  $\text{CB}[7]\cdot\mathbf{1}$  complex persisted throughout the column as well as in the transition from solution to gas phase and was detected without any fragmentation, testifying to a high association energy of  $\text{CB}[7]\cdot\mathbf{1}$ . Similar results were obtained with  $\text{CB}[7]\cdot\mathbf{2}$  (Supporting Information, Fig. S34).

### Electrochemical characterisation



Ferrocene derivatives are electrochemically active, exhibiting fast and reversible one-electron oxidation.<sup>36</sup> This offers an additional approach to examine the CB[7]-guest host-guest interactions. We investigated the electrochemical properties of the ferrocene derivatives in the absence and presence of the CB[7] host in aqueous solution using cyclic voltammetry (CV). The cyclic voltammogram of **1** shows one set of well-defined peaks corresponding to reversible one-electron transfer of the ferrocene (Fc)/ferrocenium (Fc<sup>+</sup>) couple at a midpoint redox potential ( $E^\circ$ ) of 670 mV vs. saturated calomel electrode (SCE), Fig. 4. A large anodic tail is observed at higher potential (> 900 mV vs. SCE) and is caused by the oxidation of bromide counter ions to the ammonium cations in **1**, to bromine (Supporting Information, Fig. S6).



**Fig. 4** Cyclic voltammograms of **1** (dotted line), CB[7]·**1** (dashed line), and CB[7]·**2** (continuous line). The samples (50  $\mu\text{m}$ ) are recorded using a glassy carbon electrode in aqueous  $\text{KH}_2\text{PO}_4$  (20 mM, pH 4.6). Scan rate is 50 mV/s.

The addition of CB[7] to **1** forms the inclusion complex CB[7]·**1** and leads to a pronounced 172 mV anodic shift of the ferrocene  $E^\circ$  value to 842 mV vs. SCE (Table 1). This observation is consistent with the encapsulation of the ferrocene moiety inside the CB[7] cavity.<sup>24, 28, 37, 38</sup> An excess of 0.3 equivalents CB[7] was added to **1** to ensure complete formation of the inclusion complex. A reference experiment with CB[7] alone, in the same potential window, showed no redox activity thus confirming that the peaks observed originate solely from the inclusion complex CB[7]·**1**. A large anodic tail in the CV of CB[7]·**1** at high potential (> 1.00 V vs. SCE) is again caused by oxidation of the bromide counter ions in CB[7]·**1**.

The anodic shift in  $E^\circ$  from **1** to CB[7]·**1** reflects the local environmental changes around the ferrocene moiety. The electroneutral reduced ferrocene moiety is considerably stabilized in the complex CB[7]·**1**. Upon reduction of the ferrocene centre the resulting anodic shift corresponds to a free energy drop ( $\Delta\Delta G^\circ$ ) of 16.6 kJ mol<sup>-1</sup>. This decrease is caused by stabilization of the doubly charged reduced ferrocene inside the hydrophobic cavity of CB[7] relative to the triply charged oxidized form that is less favourable for CB[7] enclosure,<sup>28, 39</sup> i.e. it becomes easier to reduce the oxidized form than for free **1**.

The large association constant of CB[7]·**1** is also evidenced by the electrochemical data. The CV of CB[7]·**1** shows only a single cathodic peak at 802 mV vs. SCE and no signal at 630 mV vs. SCE corresponding to pure **1**, indicating that the complex has not disassembled.

**Table 1** Midpoint redox potentials ( $E^\circ$ , mV vs. SCE) and potential changes on complexation of free **1** and **2** with CB[7] ( $\Delta E^\circ$ , mV vs. SCE). The associated changes in Gibbs free energy ( $\Delta\Delta G^\circ$ , kJ mol<sup>-1</sup>) are also given.<sup>[a]</sup>

Compound	$E^\circ$ [mV]	$\Delta E^\circ$ <sup>[b]</sup> [mV]	$\Delta\Delta G^\circ$ <sup>[c]</sup> [kJ mol <sup>-1</sup> ]
<b>1</b>	670	—	—
CB[7]· <b>1</b>	842	+ 172	– 16.6
CB[7]· <b>2</b>	855	+ 185	– 17.8

<sup>[a]</sup>Data obtained from cyclic voltammetry in aqueous 20 mM  $\text{KH}_2\text{PO}_4$  solution (pH 4.6), 25 °C. Scan rate 50 mV s<sup>-1</sup>. <sup>[b]</sup> $\Delta E^\circ$  values calculated as  $E^\circ(\text{CB[7]·guest}) - E^\circ(\text{1})$ . <sup>[c]</sup>Obtained using  $\Delta\Delta G^\circ = -v \cdot F \cdot \Delta E^\circ$ , where  $v$  is the number of electrons transferred in the reaction and  $F$  is the Faraday constant.

Similar CVs were observed for CB[7]·**2** with a midpoint redox potential of 855 mV vs. SCE corresponding to an anodic shift of 185 mV vs. **1**. This can be ascribed to the same reasons as for **1** and CB[7]·**1**. The fact that the complexation is clearly detected at micromolar concentrations supports the high stability of these complexes. In addition, a pair of smaller peaks is observed around 278 mV vs. SCE. This process is caused by formation of carbonyl and/or quinone compounds on the carbon electrode surface, which is common for chloride containing supporting electrolytes,<sup>40</sup> corresponding to the fact that chloride anions are the counter ion in CB[7]·**2**.

**Electrochemical kinetics for **1** and CB[7]·**1**** Cyclic voltammetry for solute **1**, CB[7]·**1** and CB[7]·**2** at shorter time scales, i.e. at faster scan rates gives voltammetric signals that display different patterns for the three compounds (Supporting Information, Fig. S7). **1** follows a diffusion controlled electrochemical process for scan rates up to 0.5 V s<sup>-1</sup>, with both the anodic ( $i_{pa}$ ) and cathodic peak current ( $i_{pc}$ ) following a scan rate ( $v$ ) dependence,  $i_{pa,c} \propto v^x$ ,  $x \approx \frac{1}{2}$  ( $R^2 = 0.996 \pm 0.001$ , see Supporting Information, Fig. S8 and S9). This roughly applies also to CB[7]·**1**, although formally with a slightly higher  $x \approx 0.6$ . For **1** the peaks separate at scan rates higher than 0.5 V s<sup>-1</sup>, where kinetic control and “quasi-reversible” behaviour gradually take over (see further discussion below and Supporting Information, Fig. S11).

The diffusion coefficients ( $D_0$ ) of the electroactive species were obtained using the Randles-Sevcik equation (Eq. 1),<sup>41</sup>

$$i_p = 0.4463 \cdot n \cdot F \cdot A \cdot c \cdot \sqrt{\frac{n \cdot F \cdot v \cdot D_0}{R \cdot T}} \quad (1)$$

where  $n$  is the number of electrons in the redox process,  $F$  is the Faraday constant,  $A$  the electrode area,  $c$  the concentration of the compound,  $R$  the gas constant, and  $T$  the absolute temperature. The resulting  $D_0$  are listed in Table 2.

Among other determinants the diffusion coefficient depends on the size (diameter) of the molecule.<sup>42</sup> As the CB[7]·**1** complex is bulkier than its free counterparts, its diffusion coefficient determined from eq. (1) is similarly smaller (Table 2) reflecting the larger hydrodynamic radius of the complex, as

seen previously for related systems.<sup>38</sup>  $D_0$  for CB[7]·1 is thus less than half of the value for **1**  $((6.3 \pm 0.7) \cdot 10^{-7} \text{ cm}^2 \text{ s}^{-1}$  vs  $(1.5 \pm 0.4) \cdot 10^{-6} \text{ cm}^2 \text{ s}^{-1}$ ), roughly following the diameter of the molecules.

**1** and CB[7]·1 thus display consistent diffusion behaviour. The calculated  $D_0$  for CB[7]·2, based on eq.(1), is, however, 16 times that for CB[7]·1 of comparable size, and six times of that for **1**. This abnormally large  $D_0$  value for CB[7]·2 almost certainly suggests a different electron transfer mechanism for CB[7]·2. While  $D_0$  for **1** and CB[7]·1 in Table 2 can be regarded as real diffusion coefficients,  $D_0$  for CB[7]·2 must therefore instead be denoted as “apparent”.

**Table 2** Diffusion coefficients ( $D_0$ ,  $\text{cm}^2 \text{ s}^{-1}$ ) and heterogeneous electron transfer rate constants ( $k_s$ ,  $\text{cm s}^{-1}$ ) obtained for **1** (real), CB[7]·1 (real) and CB[7]·2 (apparent).<sup>[a]</sup>

Compound	pH	$D_0$ [ $\text{cm}^2 \text{ s}^{-1}$ ]	$k_s$ [ $\text{cm s}^{-1}$ ]
<b>1</b>	4.6	$(1.5 \pm 0.4) \cdot 10^{-6}$	$(1.4 \pm 0.1) \cdot 10^{-2}$
CB[7]·1	4.6	$(6.3 \pm 0.7) \cdot 10^{-7}$	$(2.8 \pm 0.2) \cdot 10^{-3}$
CB[7]·2	4.6	$(9.2 \pm 0.3) \cdot 10^{-5}$	$(3.9 \pm 0.1) \cdot 10^{-2}$
	1.4	$(10.6 \pm 0.1) \cdot 10^{-5}$	$(59.2 \pm 7.9) \cdot 10^{-2}$

<sup>[a]</sup>The data derived from cyclic voltammetry in either  $\text{KH}_2\text{PO}_4$  (20 mM, pH 4.6) or  $\text{HClO}_4$  (20 mM, pH 1.4) at 25 °C.

Table 2 also summarizes the real and “apparent” heterogeneous electron transfer rate constants ( $k_s$ ) obtained for **1**, CB[7]·1, and CB[7]·2. The ferrocene moiety in **1** was found to follow reversible behaviour at low scan rates ( $< 0.5 \text{ V s}^{-1}$ ), according with previous reports for similar systems.<sup>35, 38, 43</sup> Reversible processes involve fast electron transfer compared to the scan rate with only transient surface interaction between the electrode and the reagents, and peak separation of 59 mV (for a one-electron transfer) independent of the scan rate.<sup>42</sup> Approximately this behaviour was observed for **1** as evidenced by the invariance of the peak-to-peak splitting ( $\Delta E_p = 79 \pm 1 \text{ mV}$ ) as the scan rate increased up to about  $0.5 \text{ V s}^{-1}$ . The  $\Delta E_p$  value is somewhat higher than the theoretical value of 59 mV which may be caused by small distortions due to solution resistance effects.<sup>44</sup>

When the scan rate was increased to  $4 \text{ V s}^{-1}$  the difference between the anodic and cathodic midpoint potentials of **1** was found to increase, indicating that the scan rate is now faster than the electron transfer process, and the process thus changes from reversible to quasi-reversible. The heterogeneous electron transfer rate constant,  $k_s$  for **1** was determined for the quasi-reversible process using Nicholson’s Method (Eq. 2),<sup>45</sup> and calculated at given scan rates using the parameter  $\Psi$ , which is based on the experimental  $\Delta E_p$  values, i.e.

$$\Psi = \frac{(D_0/D_R)^{\alpha/2} \cdot k_s \cdot \sqrt{R \cdot T}}{\sqrt{D_0 \cdot \pi \cdot n \cdot F \cdot v}} \quad (2)$$

where  $D_0$  and  $D_R$  are the diffusion coefficients of the oxidized and reduced species and  $\alpha$  the transfer coefficient. Taking  $D_0 = D_R$  and  $\alpha = 0.5$ ,  $k_s$  for **1** was determined for all applied scan rates and found to be  $k_s = (1.4 \pm 0.1) \cdot 10^{-2} \text{ cm s}^{-1}$  independent of the scan rate (Table 2, and Supporting Information).

The complex CB[7]·1 displays quasi-reversible behaviour at scan rates higher than  $50 \text{ mV s}^{-1}$ . The midpoint potential separation increases linearly from 79 to 325 mV as the scan rate increases from 0.05 to  $4.00 \text{ V s}^{-1}$ . The heterogeneous electron transfer rate constant was determined for CB[7]·1 similarly as for **1**, giving  $k_s = (2.8 \pm 0.2) \cdot 10^{-3} \text{ cm s}^{-1}$ .

The rate constant values in Table 2 show that encapsulation of the ferrocene moiety in **1** into the complex CB[7]·1 reduces the electrochemical rate constant by a factor of five. A similar trend was observed in related CB[7]-guest inclusion complexes.<sup>38</sup> The smaller rate constant can be understood by viewing the CB[7] macrocycle either as an additional physical barrier for electron tunnelling, or as caused by pre-organization of the CB[7]·1 complex towards more favourable electron transfer (“gated electron transfer”)<sup>46</sup> adding an additional activation energy contribution. We address this difference further in the Conclusion Section.

**Interfacial electrochemical kinetics for CB[7]·2** The mechanistic electron transfer dynamics of CB[7]·2 is different from that of **1** and CB[7]·1. The abnormal values of the “apparent” diffusion coefficient and interfacial electron transfer rate constant based on the diffusion formalism were noted. In addition the  $i_{pa,c} \propto \nu^x$  correlation accords poorly with  $x \approx 0.5$  but shows instead a  $i_{pa,c} \propto \nu^{0.8-0.9}$  behaviour (Supporting Information, Fig. S8 and S9). The diffusion formalism represented by eqs. (1) and (2) therefore does represent well the voltammetry of CB[7]·2, which accords much better with monolayer voltammetry. This is understandable if the thiolate linker group of CB[7]·2 is adsorbed more strongly on the glassy carbon electrode surface than the thiocarbamate linkers of **1** and CB[7]·1. It is not *a priori* obvious that this should be so, but the obviously different voltammetric behaviour is a strong indication that this is indeed the case.

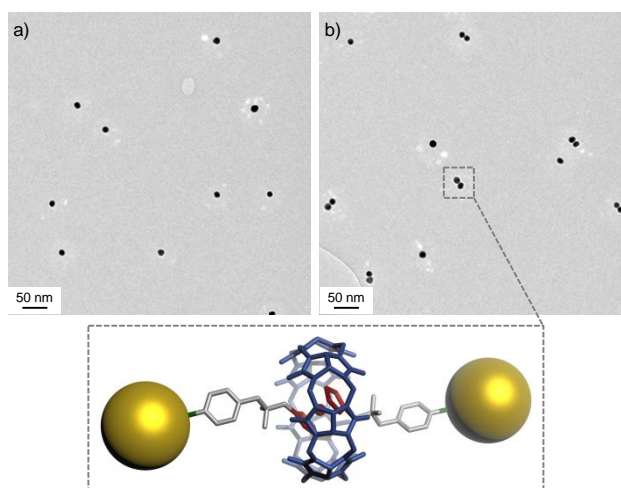
Interfacial electrochemical electron transfer kinetics for CB[7]·2 is then more appropriately based on Laviron analysis.<sup>47</sup> For the sake of completion, the Laviron analysis based standard electron transfer rate constant was calculated from the variation of the peak separation with increasing scan rate. With the reservation that the peak separations are on the small side for ideal Laviron analysis, the rate constant was found to be  $\sim 40 \text{ s}^{-1}$  (Supporting Information, Section S5.6) robust up to  $10 \text{ V s}^{-1}$  and 200 mV peak separation. This is in fact a remarkably high value testifying to highly efficient interfacial electrochemical electron transfer and strong bonding between the thiolate functionality on CB[7]·2 and the glassy carbon electrode. The different mechanisms and rate constant units mean, however, that the diffusion controlled and surface controlled rate constants ( $k_s$  in  $\text{cm s}^{-1}$  vs  $\text{s}^{-1}$ ) cannot be compared directly. In the Conclusion Section we therefore focus on the comparison of the diffusion based interfacial electrochemical rate constants of **1** and CB[7]·1 with details given in the Supporting Information, Sections S5.5-S5.7.

### Connecting gold nanoparticles

By utilizing the free thiol functionalities in CB[7]·2 it was proven possible to build a gold nanoparticle (AuNP) capped [2]rotaxane that is soluble in water. The [2]rotaxane is constructed from a

single [2]pseudorotaxane CB[7]·2 in the centre that is used as an organic linker to connect two AuNP “stoppers” by self-assembly of the thiols onto the gold surfaces. To show the general nature of the concept, the assembly was carried out using two different sizes of nanoparticles in similar experimental conditions. Notably, a solution of CB[7]·2 was exposed to either 12 or 60 nm citrate stabilized gold nanoparticles. After initial reaction, the nanoparticles dimers were purified using centrifugation (See Supporting Information, Section S6 for details).

Transmission electron microscopy (TEM) provided structural information on the [2]rotaxane “hybrid” (Fig. 5). The gold nanoparticles appear clearly as solid spheres (12 or 60 nm in diameter) while the [2]pseudorotaxane CB[7]·2 with a diameter of only ~ 1.5 nm is too small to see in the images. Without the addition of CB[7]·2 the gold nanoparticles are dispersed individually. The presence of CB[7]·2 triggers formation of AuNP dimers binding two gold nanoparticles together, and forming the desired [2]rotaxane. We note that nanoparticle rotaxane oligomers can be formed during the synthesis, but these higher density structures are removed during centrifugation.<sup>48</sup>



**Fig. 5** TEM images of the gold nanoparticle capped [2]rotaxane showing a) reference AuNP (12 nm diameter, no CB[7]·1 added) and b) after addition of CB[7]·1 with a zoom of a cartoon showing the structure of the [2]rotaxane.

## Conclusions

We have described a nanoparticle-stoppered [2]rotaxane based on an exceptionally stable pseudo-rotaxane featuring an electroactive ferrocene rod encapsulated in a cucurbit[7]uril macrocycle. The pure and nano-stoppered rotaxanes have been characterised comprehensively, and the interfacial electrochemical electron transfer patterns rationalized by a transparent frame of molecular charge transfer theory. Particularly we have been able to conclude that the electrochemical electron transfer kinetics proceeds via diffusion for 1 and CB[7]·1 while CB[7]·2 most likely operates via a highly efficient surface controlled mechanism. With these discoveries we have opened new avenues to prepare and manipulate nano-sized objects and systems, where a single or a few isolated

active molecules constitute the active units. With the enormous progress in the preparation of molecular machines seen in the past decades we believe that this work can pave the way to an entirely new type of nano-sized molecular machines.

## Conflicts of interest

There are no conflicts to declare.

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